The Practise Ground

Class 9-10 Chemistry

Types of Chemical Reactions

Four patterns you will meet again and again

1 Combination

Two or more substances form one product

$2Mg + O_2$ forms $2MgO$

2 Decomposition

One substance breaks into simpler ones

$2H_2O$ forms $2H_2 + O_2$ (electrolysis)

3 Displacement

A more reactive element replaces a less reactive one

$Fe + CuSO_4$ forms $FeSO_4 + Cu$

4 Double displacement

Two compounds swap their partners

$NaCl + AgNO_3$ forms $AgCl + NaNO_3$

Remember

Always balance the equation: the same number of each atom on both sides.

Exothermic reactions release heat. Endothermic reactions absorb it.

Balancing example

Unbalanced: $H_2 + O_2$ forms H_2O

Balanced: $2H_2 + O_2$ forms $2H_2O$

Count: 4 H and 2 O on each side

Think you've got it? Test yourself.

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