The Practise Ground

Class 11 Chemistry

Periodic Trends

How properties change across and down the table

1 Atomic radius

Decreases across a period, increases down a group

More shells down a group; stronger pull across a period

2 Ionisation enthalpy

Increases across a period, decreases down a group

Small atoms hold their electrons more tightly

3 Electronegativity

Increases across a period, decreases down a group

Ability of an atom to attract bonding electrons

4 Metallic character

Decreases across a period, increases down a group

Opposite to the trend in ionisation enthalpy

Why it happens

Across a period: nuclear charge rises while the shells stay the same, so atoms shrink.

Down a group: extra shells shield the nucleus, so atoms grow and lose electrons easily.

Quick facts

Most electronegative element: fluorine.

Noble gases have the highest ionisation enthalpy in their period.

Think you've got it? Test yourself.

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